

Article

Hydrogen Production from Palmitic Acid through Autothermal Reforming: Thermodynamic Analysis

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Abstract. This work studies thermodynamic analysis of hydrogen production via autothermal reforming of palmitic acid. A Gibbs free energy minimization method was applied to analyze thermodynamic of syngas production via oxidative reforming of palmitic acid. Equilibrium compositions were estimated at temperature of 800°C, 900°C and 1000°C under atmospheric pressure. Optimal operating temperature and molar feed ratios of steam to palmitic acid (S/C) and oxygen to palmitic acid (O/C) ranging from 0.5 to 4; were determined. The PTC Mathcad Prime 2.0 is used to calculate enthalpy, entropy and Gibbs free energy. Aspen Plus program was applied for calculation of product yields and heat duty. Maximum H₂ yield of 17.88 mole/mole_{Palmitic} with S/C of 4 and O/C of 0.5 can be achieved at 1000°C and CO yield were 7.75 mole/mole_{Palmitic} with S/C of 4 and O/C of 4. Production of hydrogen and carbon monoxide increased with increasing temperature.

Keywords: Autothermal reforming, palmitic acid, hydrogen production.

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1. Introduction

Oil crisis and shortage of fossil fuels have stimulated the exploration of alternative energy sources. Hydrogen is an attractive source of renewable energy [1] due to its high energy density by weight. Hydrogen is a combustible gas, potentially a clean energy source and industrially used as a precursor in the production of chemicals.

Hydrogen can be produced by several methods including gasification (of feedstocks such as coal, natural gas and biomass), reforming (autothermal reforming (ATR) and steam reforming (SR)), electrolysis of water and fermentation processes [2]. SR is a strongly endothermic reaction where hydrocarbons react with steam while ATR is a reaction between hydrocarbons, steam and other oxygenated compound; such as hydrogen peroxide, oxygen or air. During ATR, partial oxidation (PO) occurs when the system is stoichiometrically oxygen deficient. PO is an exothermic partial combustion of the fuel and oxygenated gas. In this process, heat released during the partial oxidation can be utilized in the steam reforming, supplying the endothermic energy requirement. This will potentially provide a thermally efficient approach when the thermoneutral condition is applied for heat integration in the ATR.

The renewable energy is one of the most promising energies for the future since it can be renewed environmentally-friendly. The exploitation of renewable fuels to fulfill our energy requirements such as hydrogen, bioethanol, dimethyl ether, and biodiesel has been widely investigated to sustain the energy requirement. Palmitic acid ($C_{16}H_{32}O_2$) is a heavy hydrocarbon obtained as a byproduct during biodiesel production. Palmitic acid can be a potential source of hydrogen. However, only a few researchers have presented the catalytic reforming or cracking of palmitic acid to hydrogen [3–7]. In this present study, the autothermal reforming processes of palmitic acid to produce hydrogen. The previous study utilized hydrogen peroxide in autothermal reformer for hydrogen production for application in micro fuel cells. Concentration of carbon monoxide in the product gas was minimized [8]. Palmitic acid steam reforming reaction contains of several minor reactions, for instance, alkanes and alkenes reforming reactions [9]. Steam was fed to reform hydrocarbons and to increase H_2 to CO ratio [10]. Oxygen was added as a co-reactant in autothermal or oxidative reforming operation, to reduce carbon formation and required external heat. However, the oxygen decreases the H_2 to CO ratio unavoidably. In this paper, the thermodynamic analysis of hydrogen production via oxidative reforming of palmitic acid using steam and oxygen is simulated. A total Gibbs free energy minimization method was adopted to estimate the equilibrium composition at temperatures of 800, 900 and 1000°C.

2. Methods

Either (1) equilibrium constant calculation method or (2) minimization of the total Gibbs free energy method should be implemented to predict thermodynamic behavior of the hydrogen production via palmitic acid autothermal reforming using oxygen. In this study, list of possible reactions; such as, alkanes and alkenes steam reforming, water-gas shift, carbon formation [11], coke gasification and Boudouard reaction (disproportionation of carbon monoxide to carbon dioxide and carbon); are summarized. Minimization of the Gibbs free energy is widely utilized in fuel reforming processes and favorable when the reaction is considered at specific temperature and pressure. To compare with the calculation of equilibrium constant method, the Gibbs free energy minimization method is simpler as the former method requires associated chemical reaction information such as enthalpy, entropy and heat capacity of each component. Therefore, the Gibbs free energy minimization method is applied in this work.

2.1. Autothermal Reforming of Palmitic Acid and Its Side Reactions

The autothermal reforming of palmitic acid has numerous undesirable side reactions because the palmitic acid is a high number of carbon atoms fatty acid. The possible reaction routes of palmitic acid autothermal reforming are shown in Table 1. Seventeen species including palmitic acid ($C_{16}H_{32}O_2$), water (H_2O), oxygen (O_2), hydrogen (H_2), carbon monoxide (CO), carbon dioxide (CO_2), methane (CH_4), alkane (C_nH_{2n+2}), alkene (C_nH_{2n}) and elemental carbon (graphite, C) as a solid were included in the simulation. Not that only gas reaction (C_1 to C_5) is considered for alkane and alkene [9].

Table 1. Reaction of autothermal reforming using palmitic acid [2].

No.	Reaction	Equation
1	Steam reforming of palmitic acid	$C_{16}H_{32}O_2 + 14H_2O \rightarrow 16CO + 30H_2$
2	Partial oxidation of palmitic acid	$C_{16}H_{32}O_2 + 15O_2 \rightarrow 16CO + 16H_2O$
3	Overall reaction of palmitic acid	$C_{16}H_{32}O_2 + 7H_2O + 7.5O_2 \rightarrow 16CO + 15H_2 + 8H_2O$
4	Methane decomposition	$CH_4 \rightarrow C + 2H_2$
5	Boudouard reaction	$2CO \rightarrow C + CO_2$
6	Water-gas shift reaction	$CO + H_2O \rightarrow CO_2 + H_2$
7	Hydrogenation of carbon monoxide(1)	$CO + 3H_2 \rightarrow CH_4 + H_2O$
8	Hydrogenation of carbon dioxide	$CO_2 + 4H_2 \rightarrow CH_4 + 2H_2O$
9	Hydrogenation of carbon monoxide(2)	$2CO + 2H_2 \rightarrow CO_2 + CH_4$
10	Hydrogenation of coke	$C + 2H_2 \rightarrow CH_4$
11	Methane steam reforming	$CH_4 + H_2O \rightarrow CO + 3H_2$
12	Methane dry reforming	$CH_4 + CO_2 \rightarrow 2CO + 2H_2$
13	Carbon monoxide reduction(1)	$CO + H_2 \rightarrow C + 3H_2O$
14	Carbon monoxide reduction(2)	$2CO + 2H_2 \rightarrow 2C + 2H_2O$
15	Reverse water-gas shift reaction	$CO_2 + H_2 \rightarrow CO + H_2O$
16	Coke gasification	$C + H_2O \rightarrow CO + H_2$
17	Oxidation of methane	$CH_4 + 2O_2 \rightarrow CO_2 + 2H_2O$
18	Partial oxidation of carbon monoxide	$CO + 0.5O_2 \rightarrow CO_2$
Alkane steam reforming		
19	Methane SR	$CH_4 + H_2O \rightarrow CO + 3H_2$
20	Ethane SR	$C_2H_6 + 2H_2O \rightarrow 2CO + 5H_2$
21	Propane SR	$C_3H_8 + 3H_2O \rightarrow 3CO + 7H_2$
22	Butane SR	$C_4H_{10} + 4H_2O \rightarrow 4CO + 9H_2$
23	Pentane SR	$C_5H_{12} + 5H_2O \rightarrow 5CO + 11H_2$
Alkane partial oxidation		
24	Metane PO	$CH_4 + 0.5O_2 \rightarrow CO + 2H_2$
25	Etane PO	$C_2H_6 + O_2 \rightarrow 2CO + 3H_2$
26	Propane PO	$C_3H_8 + 1.5O_2 \rightarrow 3CO + 4H_2$
27	Butane PO	$C_4H_{10} + 2O_2 \rightarrow 4CO + 5H_2$
28	Pentane PO	$C_5H_{12} + 2.5O_2 \rightarrow 5CO + 6H_2$
Alkene steam reforming		
29	Ethylene SR	$C_2H_4 + 2H_2O \rightarrow 2CO + 4H_2$
30	Propylene SR	$C_3H_6 + 3H_2O \rightarrow 3CO + 6H_2$
31	Butene SR	$C_4H_8 + 4H_2O \rightarrow 4CO + 8H_2$
32	Pentene SR	$C_5H_{10} + 5H_2O \rightarrow 5CO + 10H_2$
Alkene partial oxidation		
33	Ethylene PO	$C_2H_4 + O_2 \rightarrow 2CO + 2H_2$
34	Propylene PO	$C_3H_6 + 1.5O_2 \rightarrow 3CO + 3H_2$
35	Butene PO	$C_4H_8 + 2O_2 \rightarrow 4CO + 4H_2$
36	Pentene PO	$C_5H_{10} + 2.5O_2 \rightarrow 5CO + 5H_2$

No.	Reaction	Equation
Alkane oxidation		
37	Methane oxidation	$\text{CH}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$
38	Ethane oxidation	$\text{C}_2\text{H}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$
39	Propane oxidation	$\text{C}_3\text{H}_8 + 5\text{O}_2 \rightarrow 3\text{CO}_2 + 4\text{H}_2\text{O}$
40	Butane oxidation	$\text{C}_4\text{H}_{10} + 6.5\text{O}_2 \rightarrow 4\text{CO}_2 + 5\text{H}_2\text{O}$
41	Pentane oxidation	$\text{C}_5\text{H}_{12} + 8\text{O}_2 \rightarrow 5\text{CO}_2 + 6\text{H}_2\text{O}$
Decomposition		
42	Palmitic acid decomposition(1)	$\text{C}_{16}\text{H}_{32}\text{O}_2 \rightarrow 2\text{H}_2 + 7\text{CH}_4 + 7\text{C} + 2\text{CO}$
43	Palmitic acid decomposition(2)	$\text{C}_{16}\text{H}_{32}\text{O}_2 \rightarrow 2\text{H}_2 + 7\text{CH}_4 + 8\text{C} + \text{CO}_2$

2.2. Estimate Gibbs Free Energy

Gibbs free energy of each element was calculated from enthalpy, entropy and heat capacity at various temperatures as shown in Eq. (1), using PTC Mathcad Prime 2.0.

$$\Delta G = \Delta H - T \cdot \Delta S \quad (1)$$

$$\Delta H = \Delta H_f^o + \int_{T_0}^T C_{p_0} dT \quad (2)$$

$$\Delta S = \frac{1}{T_0} (\Delta H_f^o - \Delta G_f^o) + \int_{T_0}^T \frac{C_{p_0}}{T} dT \quad (3)$$

$$C_{p_0} = A + B \cdot T + C \cdot T^2 + D \cdot T^3 + E \cdot T^4 \quad (4)$$

$$\Delta G_r = \sum_{i=1}^N \gamma_i \Delta G_i \quad (5)$$

$$\Delta H_r = \sum_{i=1}^N \gamma_i \Delta H_i \quad (6)$$

$$\Delta S_r = \sum_{i=1}^N \gamma_i \Delta S_i \quad (7)$$

(i = component, $\gamma_i < 0$ for reactant component and $\gamma_i > 0$ for product component)

where ΔG is the Gibbs free energy of component in J/mol, ΔH is the enthalpy of component in J/mol, ΔS is the entropy of component in J/mol K, C_{p_0} is the capacity of component in J/g mol K, ΔG_r , ΔH_r and ΔS_r are the Gibbs free energy in J/mol, the enthalpy in J/mol and the entropy of reaction in J/mol K, respectively. A, B, C, D and E are characteristic constants of particular gas taken from Handbook of Thermodynamics Diagram Volume 1-4 [12] at each operating temperature, which were determined by Aspen Plus using RGibb reactor and SRK equation.

Equilibrium constant (K) was calculated by the following equation:

$$K = e^{\left(\frac{-\Delta G}{RT}\right)} \quad (8)$$

where K is the chemical reaction equilibrium constant for reaction, ΔG is the Gibbs free energy in kJ/mol, R is the universal gas constant in J/mol K and T is the temperature in K.

2.3. Computational Prediction

The yield and duty of each temperature were calculated from Aspen Plus program using RGibb reactor and SRK equation. Molar feed ratios of steam to palmitic acid (S/C) and oxygen to palmitic acid (O/C) are 0.5 to 4, respectively, with autothermal reaction at ambient pressure and temperatures between 800-1000°C. Methanation of CO synthesis is a catalytic exothermic process at temperatures of 473-673 K (200-400°C) and high pressure does not have any effect on hydrogen and carbon monoxide yields [13–17].

Hydrogen yield was calculated using Eq. (9)

$$Y_i = \frac{F_{iout}}{F_{P_{in}}} \quad (9)$$

where Y_i is the yield of hydrogen, F_{iout} is the molar flow rate of hydrogen at outlet, $F_{P_{in}}$ is the molar flow rate of palmitic acid at inlet.

3. Results and Discussion

3.1. Estimate Gibbs Free Energy of Reaction

PTC Mathcad Prime 2.0 program is used for component enthalpy calculation where ΔG_f and ΔH_f are the standard Gibbs free energy of formation and enthalpy of formation in J/mol, respectively [18], and temperature in °C.

Evaluation of the enthalpy, entropy and Gibbs free energy were carried out using Eq. (5) to Eq. (7), to predict the thermodynamic behavior of the palmitic acid autothermal reforming. Estimated results are shown in Table 2. Equilibrium constants were estimated using Eq. (8). Heat of reactions, difference between the total enthalpy of formation of the products and that of the reactant, were calculated to indicate whether it is endothermic or exothermic reaction [19]. For the endothermic reaction, where the heat of reaction is positive, sufficient energy from surrounding is required to drive the process forward. On the other hand, the reaction is considered to be exothermic if it shows negative heat of reaction. In this case, the internal energy needs to be released to the surrounding during the reaction. In addition, Gibbs free energy of reaction can determine the direction of reaction, for example, negative of Gibbs free energy indicates that the reaction proceeds spontaneously in the forward direction at a particular temperature and pressure while the reaction proceeds spontaneously in the reverse direction for positive of Gibbs free energy [19].

From the result, most of reactions in the autothermal process are exothermic at these three temperatures except the steam reforming of palmitic acid (R-1), overall reaction of palmitic acid (R-3), methane decomposition (R-4), methane steam reforming (R-11), methane dry reforming (R-12), reverse water-gas shift reaction (R-15), coke gasification (R-16), alkane steam reforming (R-19 to 23) and alkene steam reforming (R-29 to 32). From the Gibbs free energy, the boudouard reaction (R-5), hydrogenation of carbon monoxide (1) and (2) (R-7 and R-9), hydrogenation of carbon dioxide (R-8), hydrogenation of coke (R-10), and carbon monoxide reduction (1) and (2) (R-13 to 14) are proceeded in the reverse reaction. Interestingly, water-gas shift reaction (R-6) is spontaneously forward reaction at low temperature (800°C) while at higher temperature (900 and 1000°C), it becomes spontaneously reverse reaction [20].

Once the Gibbs free energy is obtained, the thermodynamic equilibrium constant of the reaction (K) is then calculated using Eq. (8) to estimate the extent of the reaction. When the thermodynamic equilibrium constant is larger than 1.0, it indicates that the mixture at the process outlet contains mostly products, while less than 1.0 of that states the reactants as the main components. However, if the thermodynamic equilibrium constant is equal to 1.0, it means that the reaction reaches equilibrium with the same amounts of products and reactants.

Table 3 show the spontaneous reaction and K value of three temperatures using Eq. (8). The result shows that reactive reactions are boudouard reaction (R-5), hydrogenation of carbon dioxide (R-8), carbon monoxide reduction (1) (R-7), carbon monoxide reduction (2) (R-9) and hydrogenation of coke (R-10). Water-gas shift reaction (R-6) is active at 800°C while reverse water-gas shift reaction (R-15) is active at 900 and 1000°C. Gibbs free energy of the water-gas shift reaction was increased with increasing temperature and became positive (cannot be occurred by itself) at temperature around 900°C [20].

Table 2. Enthalpy of reaction.

No.	Reaction	At 800°C		At 900°C		At 1000°C	
		ΔH (J/mol)	ΔG (J/mol)	ΔH (J/mol)	ΔG (J/mol)	ΔH (J/mol)	ΔG (J/mol)
1	Steam reforming of palmitic acid	1.87×10 ⁶	-2.31×10 ⁵	1.84×10 ⁶	-4.41×10 ⁵	1.79×10 ⁶	-6.50×10 ⁵
2	Partial oxidation of palmitic acid	-5.57×10 ⁶	-6.54×10 ⁶	-5.62×10 ⁶	-6.64×10 ⁶	-5.69×10 ⁶	-6.74×10 ⁶
3	Overall reaction of peroxide	-2.02×10 ⁶	-6.23×10 ⁶	2.00×10 ⁶	-3.63×10 ⁵	-2.01×10 ⁶	-7.02×10 ⁶
4	Methane decomposition	8.98×10 ⁴	-2.76×10 ⁴	9.04×10 ⁴	-3.86×10 ⁴	8.93×10 ⁴	-5.11×10 ⁴
5	Boudouard reaction	-1.70×10 ⁵	1.73×10 ⁴	-1.69×10 ⁵	3.46×10 ⁴	-1.70×10 ⁵	5.04×10 ⁴
6	Water- gas shift reaction	-3.45×10 ⁴	-1.10×10 ³	-3.35×10 ⁴	1.96×10 ³	-3.25×10 ⁴	4.94×10 ³
7	Hydrogenation of carbon monoxide(1)	-2.26×10 ⁵	4.60×10 ⁴	-2.26×10 ⁵	7.12×10 ⁴	-2.27×10 ⁵	9.66×10 ⁴
8	Hydrogenation of carbon dioxide	-1.91×10 ⁵	4.71×10 ⁴	-1.93×10 ⁵	6.92×10 ⁴	-1.94×10 ⁵	9.16×10 ⁴
9	Hydrogenation of carbon monoxide(2)	-2.60×10 ⁵	4.49×10 ⁴	-2.60×10 ⁵	7.32×10 ⁴	-2.59×10 ⁵	1.02×10 ⁵
10	Hydrogenation of coke	-8.98×10 ⁴	2.76×10 ⁴	-9.04×10 ⁴	3.86×10 ⁴	-8.93×10 ⁴	5.11×10 ⁴
11	Methane steam reforming	2.26×10 ⁵	-4.60×10 ⁴	2.26×10 ⁵	-7.12×10 ⁴	2.27×10 ⁵	-9.66×10 ⁴
12	Methane dry reforming	2.60×10 ⁵	-4.49×10 ⁴	2.60×10 ⁵	-7.32×10 ⁴	2.59×10 ⁵	-1.02×10 ⁵
13	Carbon monoxide reduction(1)	-1.36×10 ⁵	1.84×10 ⁴	-1.36×10 ⁵	3.26×10 ⁴	-1.37×10 ⁵	4.55×10 ⁴
14	Carbon monoxide reduction(2)	-2.72×10 ⁵	3.67×10 ⁴	-2.72×10 ⁵	6.53×10 ⁴	-2.74×10 ⁵	9.09×10 ⁴
15	Reverse water-gas shift reaction	3.45×10 ⁴	1.10×10 ³	3.35×10 ⁴	-1.96×10 ³	3.25×10 ⁴	-4.94×10 ³
16	Coke gasification	1.36×10 ⁵	-1.84×10 ⁴	1.36×10 ⁵	-3.26×10 ⁴	1.37×10 ⁵	-4.55×10 ⁴
17	Oxidation of methane	-8.02×10 ⁵	-8.01×10 ⁵	-8.03×10 ⁵	-8.02×10 ⁵	-8.03×10 ⁵	-8.02×10 ⁵
18	Partial oxidation of carbon monoxide	-2.83×10 ⁵	-1.90×10 ⁵	-2.82×10 ⁵	-1.81×10 ⁵	-2.82×10 ⁵	-1.73×10 ⁵
Alkane steam reforming							
19	Methane SR	2.26×10 ⁵	-4.60×10 ⁴	2.26×10 ⁵	-7.12×10 ⁴	2.27×10 ⁵	-9.66×10 ⁴
20	Ethane SR	2.52×10 ⁵	-2.89×10 ⁵	1.99×10 ⁵	-3.94×10 ⁵	3.74×10 ⁵	-2.68×10 ⁵
21	Propane SR	5.38×10 ⁵	-2.71×10 ⁵	5.38×10 ⁵	-3.46×10 ⁵	5.38×10 ⁵	-4.21×10 ⁵
22	Butane SR	7.10×10 ⁵	-4.23×10 ⁵	6.99×10 ⁵	-5.40×10 ⁵	6.98×10 ⁵	-6.45×10 ⁵
23	Pentane SR	8.60×10 ⁵	-5.14×10 ⁵	8.59×10 ⁵	-6.41×10 ⁵	8.58×10 ⁵	-7.69×10 ⁵
Alkane partial oxidation							
24	Methane PO	-2.25×10 ⁴	-2.35×10 ⁵	-2.26×10 ⁴	-2.54×10 ⁵	-2.28×10 ⁴	-2.74×10 ⁵
25	Ethane PO	-2.44×10 ⁵	-6.66×10 ⁵	-2.99×10 ⁵	-7.60×10 ⁵	-1.24×10 ⁵	-6.23×10 ⁵
26	Propane PO	-2.07×10 ⁵	-8.36×10 ⁵	-2.08×10 ⁵	-8.95×10 ⁵	-2.10×10 ⁵	-9.53×10 ⁵
27	Butane PO	-2.83×10 ⁵	-1.18×10 ⁶	-2.96×10 ⁵	-1.27×10 ⁶	-2.99×10 ⁵	-1.36×10 ⁶
28	Pentane PO	-3.82×10 ⁵	-1.46×10 ⁶	-3.85×10 ⁵	-1.56×10 ⁶	-3.89×10 ⁵	-1.66×10 ⁶
Alkene steam reforming							
29	Ethylene SR	2.34×10 ⁵	-1.62×10 ⁵	2.35×10 ⁵	-1.99×10 ⁵	2.35×10 ⁵	-2.36×10 ⁵
30	Propylene SR	4.08×10 ⁵	-2.54×10 ⁴	4.09×10 ⁵	-6.56×10 ⁴	4.09×10 ⁵	-1.06×10 ⁵
31	Butene SR	5.69×10 ⁵	-1.01×10 ⁵	5.69×10 ⁵	-1.63×10 ⁵	5.69×10 ⁵	-2.26×10 ⁵
32	Pentene SR	7.30×10 ⁵	-1.80×10 ⁵	7.30×10 ⁵	-2.64×10 ⁵	7.30×10 ⁵	-3.49×10 ⁵
Alkene partial oxidation							
33	Ethylene PO	-2.63×10 ⁵	-5.39×10 ⁵	-2.63×10 ⁵	-5.65×10 ⁵	-2.64×10 ⁵	-5.91×10 ⁵
34	Propylene PO	-3.36×10 ⁵	-5.91×10 ⁵	-3.38×10 ⁵	-6.15×10 ⁵	-3.39×10 ⁵	-6.39×10 ⁵
35	Butene PO	-4.24×10 ⁵	-8.55×10 ⁵	-4.26×10 ⁵	-8.96×10 ⁵	-4.29×10 ⁵	-9.36×10 ⁵
36	Pentene PO	-5.11×10 ⁵	-1.12×10 ⁶	-5.14×10 ⁵	-1.18×10 ⁶	-5.17×10 ⁵	-1.24×10 ⁶
Alkane oxidation							
37	Methane oxidation	-8.02×10 ⁵	-8.01×10 ⁵	-8.03×10 ⁵	-8.02×10 ⁵	-8.03×10 ⁵	-8.02×10 ⁵
38	Ethane oxidation	-1.55×10 ⁶	-1.61×10 ⁶	-1.61×10 ⁶	-1.67×10 ⁶	-1.44×10 ⁶	-1.50×10 ⁶
39	Propane oxidation	-2.05×10 ⁶	-2.16×10 ⁶	-2.05×10 ⁶	-2.17×10 ⁶	-2.05×10 ⁶	-2.18×10 ⁶
40	Butane oxidation	-2.65×10 ⁶	-2.88×10 ⁶	-2.67×10 ⁶	-2.91×10 ⁶	-2.67×10 ⁶	-2.93×10 ⁶
41	Pentane oxidation	-3.28×10 ⁶	-3.54×10 ⁶	-3.29×10 ⁶	-3.56×10 ⁶	-3.29×10 ⁶	-3.58×10 ⁶
Decomposition							
42	Palmitic acid decomposition(1)	-6.58×10 ⁵	-3.15×10 ⁵	-6.94×10 ⁵	-2.81×10 ⁵	-7.53×10 ⁵	-2.55×10 ⁵
43	Palmitic acid decomposition(2)	-8.28×10 ⁵	-2.98×10 ⁵	-8.63×10 ⁵	-2.47×10 ⁵	-9.22×10 ⁵	-2.05×10 ⁵

Table 3. Thermodynamic equilibrium constant for autothermal reforming reactions using palmitic acid.

No.	Reaction	Equilibrium constant (K)		
		At 800°C	At 900°C	At 1000°C
1	Steam reforming of palmitic acid	1.09	1.11	1.14
2	Partial oxidation of palmitic acid	2.05	1.95	1.96
3	Overall reaction of palmitic acid	1.50	1.47	1.49
4	Methane decomposition	1.00	1.00	1.01
5	Boudouard reaction	*	*	*
6	Water- gas shift reaction	1.00	*	*
7	Hydrogenation of carbon monoxide(1)	*	*	*
8	Hydrogenation of carbon dioxide	*	*	*
9	Hydrogenation of carbon monoxide(2)	*	*	*
10	Hydrogenation of coke	*	*	*
11	Methane steam reforming	1.01	1.01	1.01
12	Methane dry reforming	1.01	1.01	1.01
13	Carbon monoxide reduction(1)	*	*	*
14	Carbon monoxide reduction(2)	*	*	*
15	Reverse water-gas shift reaction	*	1.00	1.00
16	Coke gasification	1.00	1.00	1.00
17	Oxidation of methane	1.09	1.09	1.09
18	Partial oxidation of carbon monoxide	1.02	1.02	1.02
Alkane steam reforming				
19	Methane SR	1.01	1.01	1.01
20	Ethane SR	1.03	1.04	1.03
21	Propane SR	1.03	1.04	1.04
22	Butane SR	1.05	1.06	1.07
23	Pentane SR	1.06	1.07	1.08
Alkane partial oxidation				
24	Metane PO	1.03	1.03	1.03
25	Etane PO	1.08	1.08	1.07
26	Propane PO	1.10	1.10	1.10
27	Butane PO	1.14	1.14	1.15
28	Pentane PO	1.18	1.17	1.19
Alkene steam reforming				
29	Ethylene SR	1.03	1.02	1.02
30	Propylene SR	1.08	1.01	1.01
31	Butene SR	1.01	1.02	1.02
32	Pentene SR	1.02	1.03	1.04
Alkene partial oxidation				
33	Ethylene PO	1.06	1.06	1.06
34	Propylene PO	1.07	1.07	1.07
35	Butene PO	1.01	1.10	1.10
36	Pentene PO	1.13	1.13	1.14
Alkane oxidation				
37	Methane oxidation	1.09	1.09	1.09
38	Ethane oxidation	1.20	1.21	1.18
39	Propane oxidation	1.27	1.28	1.28
40	Butane oxidation	1.38	1.39	1.39
41	Pentane oxidation	1.49	1.49	1.49
Decomposition				
42	Palmitic acid decomposition(1)	1.04	1.03	1.03
43	Palmitic acid decomposition(2)	1.03	1.03	1.02

* K cannot be occurred by itself.

3.2. Heat Duty Calculation

The heat duty and yield of hydrogen from steam reforming, partial oxidation reforming and autothermal reforming of palmitic acid using oxygen were determined at three different temperatures, 800, 900 and 1000°C, under atmospheric pressure using Aspen Plus program where the reactions at each temperature are defined by the minimum Gibbs free energy method. Figure 1 shows heat duties (watts) of autothermal reforming of palmitic acid where the yield of hydrogen is illustrated in Fig. 2.

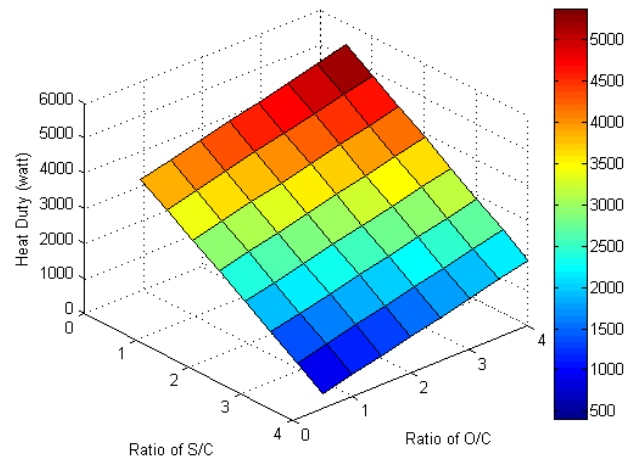
For all range of S/C ratio (varied from 0.5 to 4), the heat duties of palmitic acid steam reforming decreases with increasing temperature. When the operating temperature was constant, the heat duties show a increasing trend when the ratio S/C is increased because the higher S/C promotes the endothermic reactions such as steam reforming of palmitic acid (R-1), overall reaction of palmitic acid (R-3), methane decomposition (R-4), methane steam reforming (R-11), methane dry reforming (R-12), coke gasification (R-16), alkane steam reforming (R-19 to 23) and alkene steam reforming (R-29 to 32); and prohibit the exothermic reactions such as reaction partial oxidation of palmitic acid (R-2), water-gas shift reaction (R-6), oxidation of methane (R-17), partial oxidation of carbon monoxide (R-18), alkane partial oxidation (R-24 to 28), Alkene partial oxidation (R-33 to 36), alkane oxidation (R-37 to 41) and decomposition of palmitic acid (R-42 and 43).

From the results, higher S/C ratio and lower temperature are favored. The optimum condition was at S/C ratio of 3 at 800°C. Figure 1 shows that the net heat duties of autothermal reforming of palmitic acid. From the graph, the thermoneutral operation conditions can be obtained for S/C of 0.5 and O/C of 4 at 800°C (408.86 watt with lowest hydrogen yield of 14.94 mole/mole_{Palmitic}).

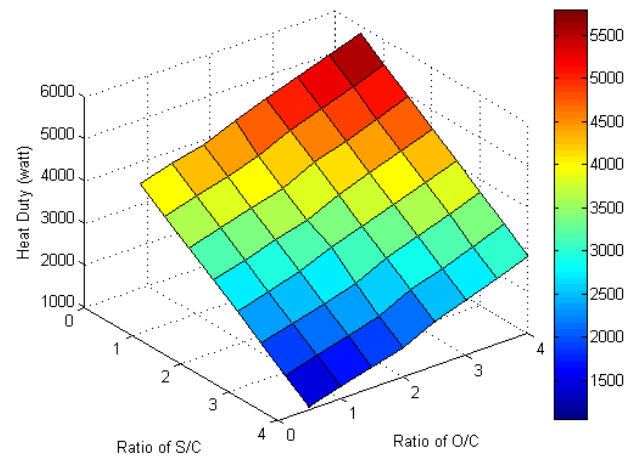
The highest endothermic of autothermal reforming using palmitic acid at ratio for S/C of 4 and O/C of 0.5 at 1000°C (5,856.78 watt) can produce a maximum hydrogen yield of 17.88 mole/mole_{Palmitic} as shown in Figs. 1 and 2.

The thermodynamic analysis predicted that, during palmitic acid autothermal reforming, hydrogen can be produced from steam reforming of palmitic acid (R-1) in Table 1, overall reaction of palmitic acid (R-3), methane decomposition (R-4), water-gas shift reaction (R-6), methane steam reforming (R-11), methane dry reforming (R-12), coke gasification (R-16), alkane steam reforming (R-19 to 23), alkane partial oxidation (R-24 to 28), alkene steam reforming (R-29 to 32), alkene partial oxidation (R-33 to 36), alkane oxidation (R-37 to 41) and decomposition of palmitic acid (R-42 and 43), whereas hydrogen can be utilized in the reverse water-gas shift reaction (R-15). Figure 2 shows the hydrogen yield as a function of S/C ratio and O/C ratio at different temperatures.

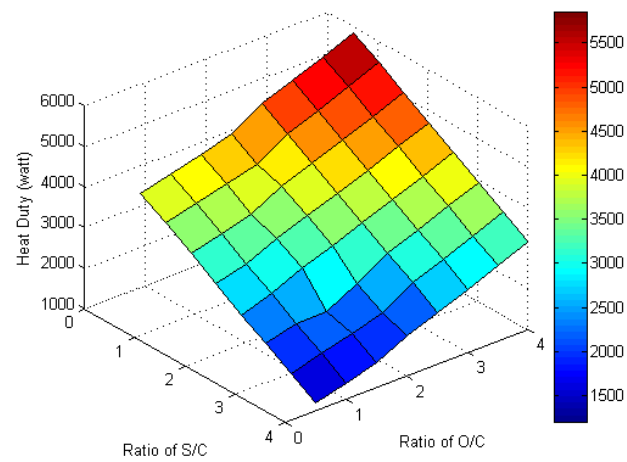
In the palmitic acid autothermal process, carbon monoxide can be produced by steam reforming of palmitic acid (R-1), partial oxidation of palmitic acid (R-2), overall reaction of palmitic acid (R-3), methane steam reforming (R-11), methane dry reforming (R-12), reverse water-gas shift reaction (R-15), coke gasification (R-16), alkane steam reforming (R-19 to 23), alkane partial oxidation (R-24 to 28) and alkene steam reforming (R-29 to 32), alkene partial oxidation (R-33 to 36) and decomposition of palmitic acid (R-42 and 43), while it can be consumed in the water-gas shift reaction (R-6) and partial oxidation of carbon (R-18). The carbon monoxide yield as a function of S/C ratio and O/C ratio at different temperatures (Fig. 3) can be observed that the yield is increased with increasing temperature as well as increasing both S/C and O/C ratios. The maximum yield of 7.75 mole/mole_{Palmitic} and minimum yield of 2.43 mole/mole_{Palmitic} can be obtained at S/C of 4 and O/C of 4 at 1000°C and at S/C of 0.5 and O/C of 0.5 at 800°C, respectively.



(a)

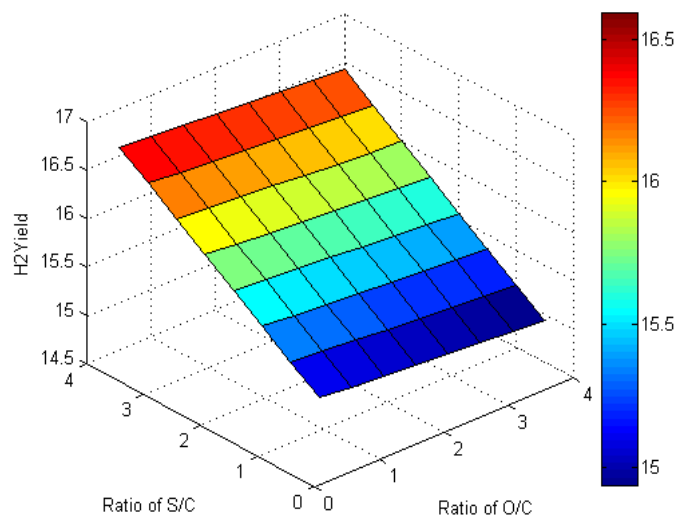


(b)

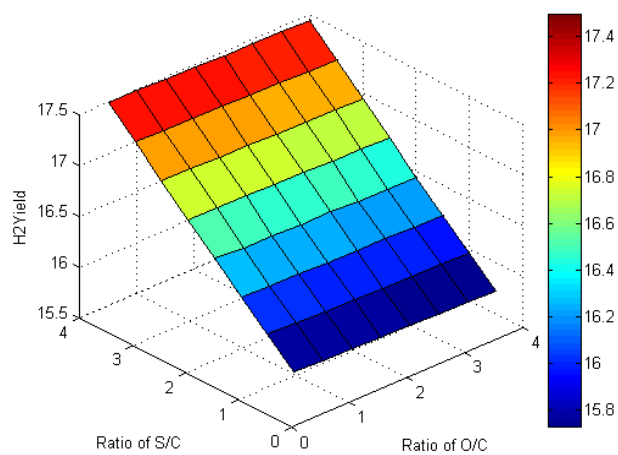


(c)

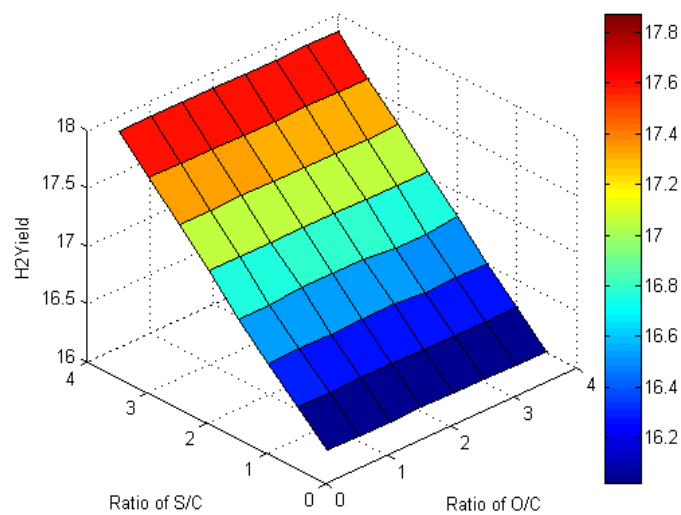
Fig. 1. Heat duty of autothermal reaction as a function of S/C ratio and O/C ratio at 1 atm and temperature: (a) 800°C, (b) 900°C and (c) 1000°C.



(a)

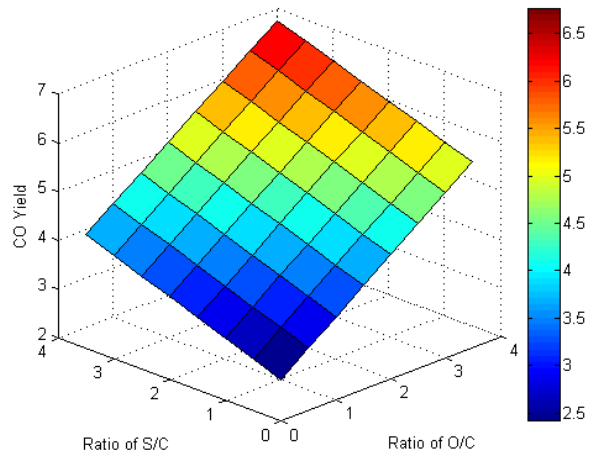


(b)

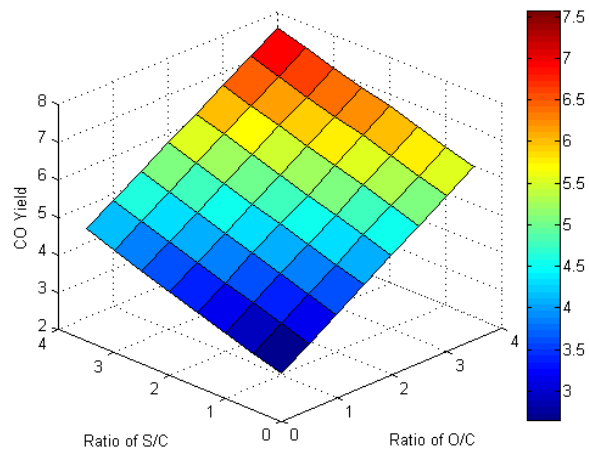


(c)

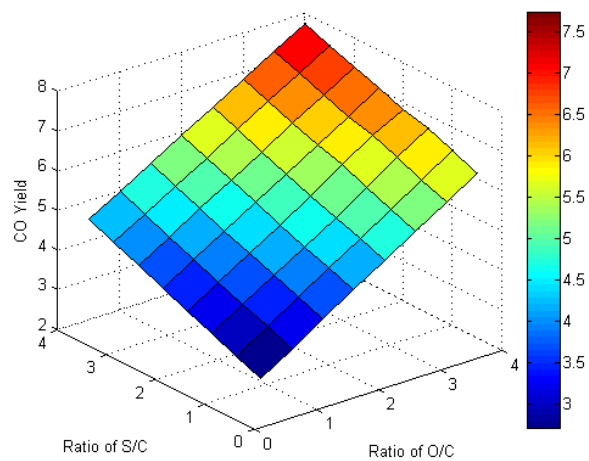
Fig. 2. Hydrogen yield as a function of S/C ratio and O/C ratio at 1 atm and temperature: (a) 800°C, (b) 900°C and (c) 1000°C.



(a)



(b)



(c)

Fig. 3. Carbon monoxide yield as a function of S/C ratio and O/C ratio at 1 atm and temperature: (a) 800°C, (b) 900°C and (c) 1000°C.

4. Conclusions

Source of hydrogen production can be fossil fuel, higher hydrocarbon and also biomass. Since palmitic acid has high carbon atom, it is suitable to be the reactant for hydrogen production. In this study, the thermodynamic analysis was performed for all possible reactions at each temperature using minimization of the total Gibbs free energy. The net heat duties of steam reforming reaction and partial oxidation reaction decrease when increasing of water ratio and oxygen ratio, respectively. The selected operating temperature strongly influenced the hydrogen yields as well as the heat duty of reaction. Hydrogen production by steam reforming of hydrocarbons, partial oxidation of hydrocarbons and water-gas shift reaction are thermodynamically favored at high temperatures, resulting in higher yields of hydrogen at higher operating temperatures. At low temperatures, hydrogen yield was relatively low due to the formation of alkane by reaction of hydrogen and carbon monoxide. The highest endothermic at ratio for S/C of 4 and O/C of 0.5 with highest yield of hydrogen was found to be 17.88 mole/mole_{Palmitic} at 1000°C.

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